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The periodic law states that the properties of elements are periodic functions of their atomic numbers. This means that the \_\_\_\_\_ determines the position of each element in the periodic table. This means that the \_\_\_\_\_ determines the position of each element in the periodic table.

### **CHAPTER 5 REVIEW THE PERIODIC LAW SECTION 1**

CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided.

1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered.

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## Chapter 5 Review The Periodic Law Answers 5 1

CHAPTER 5 REVIEW \ The Periodic Law . SHORT ANSWER Answer the following questions in the space provided. 1. Consider the neutral atom with 53 protons and 74 neutrons to answer the following questions. 53. a. What is its atomic number? 127. b. What is its mass number? atomic number. c. Is the element's position in a modern periodic table ...

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The left side of the periodic table is composed of metals, which have few valence electrons. Moving to the right, there are more electrons and their mobility is less, reducing their metallic properties. When most of the orbitals are filled, the atoms take on nonmetallic characteristics. When orbitals are full, a noble gas configuration is attained. Metals become positive ions. When sodium ...

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